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Questions pH

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Answers~~

Solution: The reaction of the acid + base gives a product of salt + water, which is considered as neutralization reaction.

Examples: $\text{NaOH} + \text{HCl} \rightarrow \text{NaCl} + \text{H}_2\text{O}$

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2 Acids Bases
and Salts with
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Solutions. Acids
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Extra Questions
Very Short
Answer Type.

Question 1. Two
solutions A and
B have pH values
of 5 and 8
respectively.
Which solution

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will be basic in
nature . . .

Salts Answers

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Extra Questions
with . . .~~

Ans:- Salts are
classified into
four categories
on the basis of
the strength of
acid and base.
they are: Salt

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of strong acid
and strong
base. e.g. NaCl,
KCl; Salt of
strong acid and
weak base. e.g.
NH₄Cl, NH₄I;
Salt of weak
acid and strong
base. e.g. CH₃
COONa, COO⁻Na
+ Salt of weak
acid and weak
base. e.g. CH₃

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$\text{COO}^- \text{NH}_4^+$ How
is common salt
found in nature?

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2. Bases react

with . acids to

produce salts

and water; water

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~~to produce acids
and salts; salts
to produce acids
and water;
neither acids,
salts, nor water~~

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Choice Questions
and Answers~~

The salts of
strong acids and
weak bases give

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Basic solution
having pH less
than 7. Example,

NH_4Cl ,

Ammonium

Chloride will

have pH less

than 7. The

salts of weak

acids and strong

bases give basic

solution having

pH more than 7.

Example, Na_2CO_3

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3, Sodium
Carbonate will
have pH more
than 7. Question
7.

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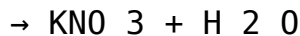
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Answer A

reaction in which an acid and base react with each other to give a salt and water is termed as neutralization reaction. For Example: (i)

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15. Give two important uses of washing soda and baking soda.

Answer Two important uses of washing soda are:

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Answer. (a) HCl is acid and NaOH is base whose combination forms the common salt. Its formula is NaCl (Sodium chloride). It is obtained from sea water. (b)

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Rock salt is the common name for the mineral "halite". Its chemical formula is NaCl . It may be white or light blue or yellow depending upon impurities present in it.

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aniline as an

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(A) inorganic acid (B) organic acid (C) soft organic acid (D) strong inorganic acid. Answer.

Answer: (a) 5.

When milk of magnesia reacts with acetic acid it produces (A)

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basic salts (B)
acidic salts (C)
neutral salt (D)
complex salt.

Answer. Answer:

(c) 6. Which of the following phenomena will occur when a small amount of acid is added to water ...

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fine as soon as

knowing the

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Turmeric is a natural indicator. On adding its paste to acid and base separately, which colours would be observed. (a) Yellow in both

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Bases And base

(b) Yellow in acid and red in base (c) Pink in acid and yellow in base (d) Red in acid and blue in base. Answer.

Answer: (b)

Yellow in acid and red in base

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Answers to help
students

understand the
concept very
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Answers. Choose
the correct
answer: Question
1.

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Acids, Bases and Salts In this first quiz on GCSE Chemistry we take a look at one vital part of the subject - acids, alkalis, bases

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and salts. In it we find out the difference between alkalis and bases, how acids and alkalis react with one another, how neutralisation can produce salts and the relevance of pH values.

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~~| Revise pH~~

~~Values and~~

~~Alkalis~~

Acids, bases and salts An acid is a compound that is defined by its physical and chemical properties.

Acids taste sour

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and react with metals and polyatomic ions called carbonates. A carbonate is a charged cluster of Carbon and Oxygen atoms.

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Acid + Base ---->
Salt + Water.

Question 3: Rain
becomes acidic
due to. carbon
dioxide. sulphur
dioxide.

nitrogen
dioxide. all of
the above.

Correct Option

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is : 4. Solution

: Rain becomes acidic because of carbon dioxide, sulphur dioxide and nitrogen dioxide. They dissolve in rain drops to form carbonic acid, sulphuric acid and nitric acid

...

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