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4.2 The Structure
of an Atom We have
learnt Dalton ' s
atomic theory in
Chapter 3, which
suggested that the
atom was indivisible
and indestructible.
But the discovery

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of two fundamental particles (electrons and protons) inside the atom, led to the failure of this aspect of Dalton ' s atomic theory. It was then considered necessary to

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Terms in this set
(34) What is an atom? the smallest particle on an element that retains its identity in a chemical reaction.

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Atom. 100. The

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Structure of the
Atom. Graphite
surface.

BIG Idea Atoms are the fundamental building blocks of matter. 4.1 Early Ideas About Matter.

MAIN Idea The ancient Greeks tried to explain matter, but the scientific study of the atom began with

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John Dalton in the
early 1800s.

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Structure of the
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1. All elements are
composed of tiny
invisible particles
(atoms) 2. All
atoms of the same
element are

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identical. 3. Atoms of different elements can physically combine together or chemically combine in simple whole number ratios to form compounds.

4. Chemical reactions occur when atoms are separated, joined, or rearranged.

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Answer Directions:

Read all questions carefully and

answer to the best of your ability. Do

not leave any questions blank!

Show your work where needed 1.

How long ago did Democritus propose the idea of atoma?

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STRUCTURE OF
THE ATOM CLASS

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SADRAS ,

KALPAKKAM 2.

1a) Charged

particles in matter

:- Atoms have three
types of sub atomic

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particles. They are electrons, protons and neutrons.

Electrons are negatively charged (e^-), protons are positively charged (p^+) and neutrons have no charge (n).

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refers to the
structure of atom

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comprising a nucleus (center) in which the protons (positively charged) and neutrons (neutral) are present. The negatively charged particles called electrons revolve around the center of the nucleus.. The history of atomic structure and

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quantum mechanics
dates back to the
times of

Democritus, the
man who first
proposed that
matter is composed
of ...

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Electrons, Protons,
Neutrons and
Atomic ...~~

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Atomic number 4,
mass number 9. It
has 4 protons, 4
electrons and $9 - 4$
 $= 5$ neutrons

Isotopes Atoms of
the same element
can have different
numbers of
neutrons; these
atoms are called
isotopes of that
element. Isotopes
of an element have

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the same chemical properties because they have the same electronic structure

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Number, Isotopes
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Atomic Number,

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Plasma Membrane

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(4.1) Dalton's atomic theory 1. all elements are made of tiny indivisible particles called atoms 2. atoms of the same element are identical and those of different

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elements are
different 3. atoms
of different
elements can
physically mix or
chemically combine
together in whole
number ratios to
form compounds

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