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To find molarity, divide the number of moles by the volume in dm³. 0.0175 mol / 0.25 dm³. = 0.07 M. So initially, solution EF was prepared so that the molarity of oxalic acid was 0.07 M, but we know from question (b) that the molarity obtained in the titration is only 0.028 M.

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The formula is: $M_1 V_1 = M_2 V_2$ Where M_1 and V_1 are the molarity and volume before a dilution, and M_2 and V_2 are the molarity and volume after a dilution. Consider this example from the 2006 NECTA: "Hydrochloric acid solution made by diluting 750 cm³ of 0.25 M HCl to 937.5 cm³ with distilled water."

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