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368 Chapter 11 □ Stoichiometry Section
111.11.1 Objectives Describe the types of

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relationships indicated by a balanced chemical equation. State the mole ratios from a balanced chemical equation.

Review Vocabulary reactant: the starting substance in a chemical reaction

New Vocabulary stoichiometry mole ratio

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Section 11.1 What is stoichiometry? 1. true 2. true 3. false 4. true 5. true 6. 2, 2, 64.10 7. 3, 3, 96.00 8. 2, 2, 88.02 9. 4, 4, 72.08 10. methanol and oxygen gas 11. carbon dioxide and water 12. 160.10 g 13. 160.10 g 14. They are equal. 15. A mole ratio is a ratio between the numbers of moles

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Stoichiometry Key Section 11.1 What is stoichiometry? In your textbook, read about stoichiometry and the balanced equation. For each statement below, write true or false. 2. 3. 4. 5. The study of the quantitative relationships between the amounts of reactants used and the amounts of products formed by a chemical reaction

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Stoichiometry Stoichiometry CHAPTER
11 SOLUTIONS MANUAL Section 11.1
Defining Stoichiometry pages 368□372

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Practice Problems pages 371–372 1.

Interpret the following balanced chemical equations in terms of particles, moles, and mass. Show that the law of conservation of mass is

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11.2 Stoichiometric Calculations Concepts
Mole-Mole ...

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Angelina_Clapp. Chapter 11
Stoichiometry. stoichiometry. mole ratio.
excess reactant. limiting reactant. The

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study of quantitative relationships between the amounts of \square . In a balanced equation, the ratio between the numbers of moles \square . A reactant that remains after a chemical reaction stops.

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In Section 11.3, for example, you learned how to express the stoichiometry of the reaction for the ammonium dichromate volcano in terms of the atoms, ions, or molecules involved and the numbers of moles, grams, and formula units of each (recognizing, for instance, that 1 mol of ammonium dichromate produces 4 mol of

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water). This section describes how to use the stoichiometry of a reaction to answer questions like the following: How much oxygen is needed to ensure complete combustion of a ...

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11.1 Defining Stoichiometry MAIN Idea

The amount of each reactant present at the
start of a chemical reaction determines

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Answer Key
how much product can form. 11.2

Stoichiometric Calculations MAIN Idea

The solution to every stoichiometric problem requires a balanced chemical

equation. 11.3 Limiting Reactants MAIN

Idea A chemical reaction Chapter 11:

Stoichiometry Stoichiometry Section 11.1

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